



**UNIVERSITI PUTRA MALAYSIA**

***DESIGN OF A DEVICE TO SAMPLE AMMONIA GAS EMISSION FROM  
FLOODED RICE SYSTEMS***

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FROM FLOODED RICE SYSTEMS**

**BY**

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**July 2019**

## APPROVAL SHEET

This project report here to entitle “**DESIGN OF A DEVICE TO SAMPLE AMMONIA GAS EMISSION FROM FLOODED RICE SYSTEMS**” was prepared and submitted by in partial fulfilment of the requirement for the degree of Bachelor of Engineering (Agricultural and Biosystem) is hereby accepted.

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## ABSTRACT

Ammonia gas volatilization in agriculture sector mainly in flooded rice systems needs to be studied in order to increase yield and productivity and reduce losses mainly in terms of economic. However, there is no research that have been conducted to measure how much ammonia gas loss from Nitrogen containing fertilizers applied in flooded rice fields in Malaysia. Monitoring the ambient ammonia concentration at paddy fields and its surrounding is important because in Malaysia, residential area are located nearby. Excessive ammonia exposure at high concentration and long period of time can cause burning of the nose, throat and respiratory tract meanwhile for low concentrations can cause coughing, and nose and throat irritation. Readily available ammonia detector and sensors are sold at high price with few limitations such as detection limit where it cannot detect ammonia gas at a low concentrations since ammonia levels in the natural atmosphere can be very low due to wind and temperature. Monitoring atmospheric ammonia concentrations does not require highly rapid detectors. The gas analysis setup also should include a selective mechanism that only significant gas can affect the detector's surrounding system. In this study, a simple device to sample the ammonia gas at a low concentrations in flooded rice system is designed.

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## CHAPTER 1

### INTRODUCTION

#### 1.1 Overview

Nitrogen fertilizers are necessary for the rice crop growth and in increasing the yield. It promotes rapid growth of plants and improves the yield and quality of grains. Stunted plants with small yellowish green leaves and fewer tillers are symptoms of nitrogen deficiency. In agricultural systems, the application of fertilizers that contain nitrogen are very broadly used. Losses of ammonia volatilization in flooded soils range from insignificant to nearly 60% of the nitrogen used (Choudhury & Kennedy, 2005).

In flooded rice systems, ammonia volatilization occurs at the surface of the flood water where nitrogen containing fertilizers are applied. As urea  $\text{CO}(\text{NH}_2)_2$  is easily hydrolysed to ammonium carbonate  $((\text{NH}_4)_2\text{CO}_3)$  in water, the carbon dioxide ( $\text{CO}_2$ ) produced starts to convert to bicarbonate ion ( $\text{HCO}_3^-$ ) resulted in increasing value of pH. High pH value in soil is one of the factors of increasing ammonia volatilization.

The use of carbon dioxide by algae and other aquatic life for their photosynthetic operations increases floodwater pH, resulting in significant losses of N due to volatilization of ammonia. Ammonium ions are weakly bonded to water molecules and converted into non-ionized ammonia ( $\text{NH}_3$ ), which can leave water as gas.

Another causes of increasing ammonia volatilization includes low rainfall frequency, moist soil or soil with high moisture content, high wind velocity, high soil temperature, low cation exchange capacity soil and poor buffering of soil.

Ammonia when reacts with sulphate or nitrate could produce aerosol particles. This aerosol does not absorb sunlight but reflect it, thus reducing the amount of sunlight reaching the surface of the Earth. Ammonia is extremely toxic to aquatic organisms, and its concentration in U.S. water has been regulated to a maximum of 0.02 mg/L (U.S. Environmental Protection Agency, 1977).

## **1.2 Problem statements**

The study of ammonia volatilization in agriculture has been conducted around the world especially in the United States, Vietnam, China and other agriculture practicing countries. The study of ammonia volatilization in flooded rice systems need to be conducted in Malaysia since this country yield about 2.5 metric ton per hectare of rice in 2016. Comparing to other ASEAN countries such as Thailand, Indonesia, Philippines and Vietnam, Malaysian rice production remained relatively constant in comparison with other countries, which showed a growing trend since 1990 (Omar, Shaharudin, & Tumin, 2019).

The volatilization of ammonia from the utilization of nitrogen based fertilizers can cause much losses to farmers. If the amount of volatilization is up to 60% of the N applied on the field, there would be such a terrible waste of cost them because only few of the remaining N applied will be utilized by the crop, not accounting other N losses by different paths.

A simple device needed to be created for the farmers so that they can keep track of how much N are dissipated through ammonia volatilization.

### 1.3 Objectives

The main objective of this project is to design a device to sample ammonia gas in flooded rice systems. The specific objectives of this project are :

- I. To study and compare the existing methods for measuring ammonia volatilization.
- II. To design a device for sampling ammonia gas in the air.
- III. To fabricate the designed device.



## CHAPTER 2

### LITERATURE REVIEW

#### 2.1 Mechanism of ammonia volatilization

Ammonia (NH<sub>3</sub>) volatilization takes place roughly two or three weeks after fertilizer use, as previous research suggested (Bouwman et al., 2002; Zhang et al., 2005; Su et al., 2007; Li et al., 2008b,c). This reaction depends on the hydrolysis of the urea itself. We can see the mechanism of ammonia volatilization through these chemical reactions :

Equation 1 :



Equation 2 :



Equation 3 :



In equation 1, the hydrolysis process occurs when the urea is applied on a moist soil or soil with high humidity with the help of urease enzymes that found naturally in the soil. Moving on to equation 2, hydroxide produced increase the soil pH temporarily and increase the rate of ammonia volatilization.

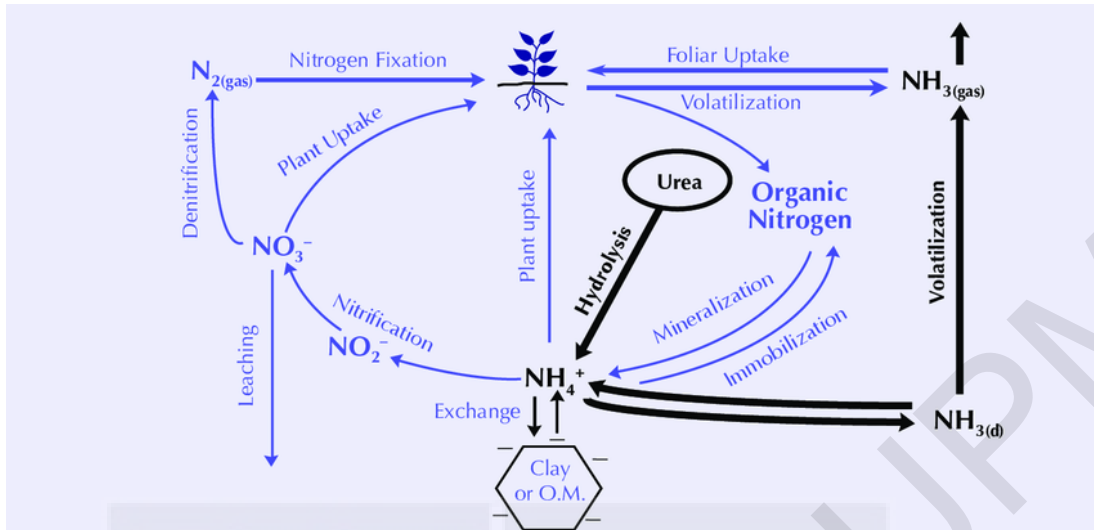


Figure 1 : Nitrogen cycle, with emphasis on volatilization of ammonia (Clain et al., 2007)

The type of nitrogen fertilizer was found to be the major factor responsible for loss of ammonia volatilization in direct seeding paddy fields. Higher levels of ammonia volatilization have been observed in Eastern China, especially in the Taihu Lake region's paddy fields (Li, Y., Huang, L., Zhang, H., Wang, M., & Liang, Z., 2017).

## 2.2 Mechanism of ammonia volatilization in flooded rice fields

In flooded rice systems, the floodwater zone is where ammonia volatilization occurred.. Ammonium containing fertilizers dissociate straight from  $\text{NH}_4$ , while urea may break down to generate  $\text{NH}_4^+$  ions by catalytic hydrolysis. When  $\text{NH}_4^+$  is transformed to  $\text{NH}_3$  gas, volatilization can cause it to be lost into the ambience. This is a significant consequence of N loss for rice fields that have been flooded with water.

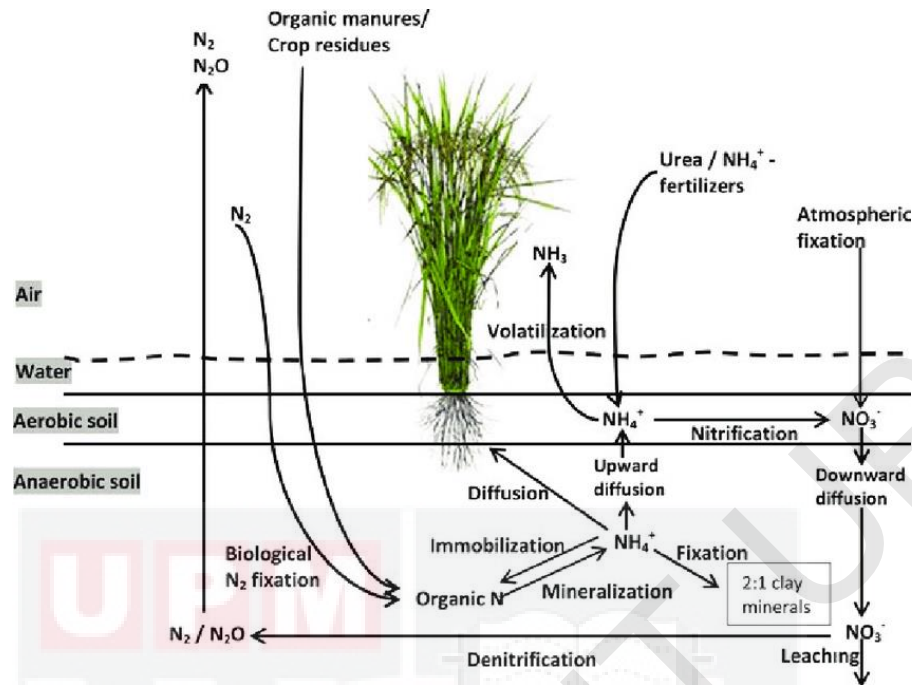


Figure 2 : Nitrogen transformation in submerged soil (Chauhan, Jabran & Mahajan, 2017)

### 2.3 Current status of ammonia volatilization

It has been brought into being that, in the worldwide, up to 64% (a standard of 18%) of utilized N was lost as  $\text{NH}_3$ . As a species of reactive Nitrogen, ammonia ( $\text{NH}_3$ ) is lost via volatilization as one of the main pathways of N loss in farming systems. While more than 40% of the N applied is described as being lost as  $\text{NH}_3$  under certain ambient and edaphic conditions (Singh et al., 2013), an average of 14% of N is lost via volatilization from artificial fertilizers (Bouwman et al., 2002, De Klein et al., 2006). In China, the total  $\text{NH}_3$  discharge was reckoned to be 13.6 Tg for 2000, of which 50% comes from fertilizer applications and another 38% from other sources of agriculture (Streets D. G. , 2003).

Asia's NH<sub>3</sub> emissions are 19.3 Tg N·yr<sup>-1</sup>, 46% of the world's total. The main source is volatilization from livestock manufacturing and rural human waste, which accounts for 62% of total emissions, and the residual 38% is from nitrogen fertilizer implementation (Zhu, Xiong, & Xing, 2005).

Table 1 : Estimation of Nitrogen (N) flux from Asia in 1999/2001 and 2030 (Zhu et al., 2005)

	Year	East Asia	South East Asia	South Asia	Central Asia	West Asia	Asia Total
<b>NH<sub>3</sub> emission/Tg N·yr<sup>-1</sup></b>	1999/2001	8.9	1.9	6.9	0.27	1.4	19.3
	2030	11.2	2.4	8.8	0.41	2.2	25.0

Table 2 : Ammonia emission in Asia countries based on emission sources (Zhu et al., 2005).

	Year	Total emissions/Tg N·yr <sup>-1</sup>	Source (in % of the total)		
			Fertilizer N	Animal excreta	Human excreta
<b>NH<sub>3</sub> emissions</b>	1999/2001	19.3	38	53.7	8.4
	2003	25	40.3	53.6	6.4

Table 3 : Review of concentrations of observed NH<sub>3</sub> at different locations. (Wang, S. et al., 2015)

<b>Placement</b>	<b>Type</b>	<b>Span</b>	<b>NH<sub>3</sub> (ppb)</b>	<b>Approach</b>
<b>Shanghai</b>	Urban	2013.7–2014.9	6.2 ± 4.6	DOAS
<b>China</b>	Rural	2013.7–12, 2014.3–6	12.4 ± 9.1	MARGA
	Industrial	2014.1–6	17.6 ± 9	DOAS
<b>Beijing</b>	Urban	2007.1.23–2.14	7.21 ± 4.941	Annular Denuder
<b>China</b>		2007.8.2–31	33.46 ± 9.111	
<b>Beijing</b>	Urban	2008.2–2010.7	22.8 ± 16.3	Passive Sampler
<b>China</b>	Rural	2007.1–2010.7	10.2 ± 10.8	
<b>North Plain</b>	Rural sites	2008.8–2009.9	20.61	Passive Sampler
<b>Kampur</b>	Urban	2007.4.8–6.30	23.7 ± 5.11	Online NO <sub>x</sub> –NH <sub>3</sub>
<b>India</b>		2007.12.1– 2008.1.31	21.5 ± 6.61	
<b>Seoul</b>	Urban/GJ	2010.9.1– 2011.8.23	10.9 ± 4.25	WS-CRDS2
<b>Korea</b>	Urban/GS		12.3 ± 4.23	
<b>Lahore</b>	Urban	2005.12– 2006.2	30.3–116.9	Annular Denuder
<b>Taiwan</b>	Industrial	2003.9– 2004.12	100.2	Passive Sampler
<b>USA</b>	Urban/Atlanta	2007.7–12	1.35 ± 1.19	Citric Acid Denuder

	Rural/Georgia		$3.32 \pm 2.37$	Difference Technique
<b>Houston, TX</b>	Urban	2010.2.12–3.1	$2.42 \pm 1.16$	EC-QCL-based
<b>USA</b>		2010.8.5–9.25	$3.07 \pm 2.87$	sensor3
<b>Wisconsin</b>	Urban	2009.1.1–3.31	2.3	iCAMs4
<b>USA</b>	Rural		2.4	
<b>USA</b>	Forest/Brent	2013.6.1–7.15	1–2	CIMS5
	Urban/Kent	2013.8.31–9.20	Up to 6	
<b>Ontario</b>	Rural	2010.3.30– 2011.3.29	4.71	Passive Sampler
<b>Vredoped</b>	Rural	2009.12.16– 2010.2.18	Up to 197.61	DOAS
<b>Barcelona</b>	Urban BC	2011.5.6–9.7	$2.9 \pm 1.3$	On-line Instrument
<b>Spain</b>	Urban CC	2011.5.13–6.28	$7.5 \pm 2.8$	

#### 2.4 Factors increasing ammonia volatilization

There are many environmental factors that can trigger ammonia volatilization including high pH value and high temperature of the soil. High soil pH and high temperature of the soil increase accumulations of ammonia dissolved in soil water and warm soil water cannot sustain as much ammonia gas. Although a temporary increase in pH may result in significant volatilization loss from soils with an initial pH as low as 5.5, warm temperatures raise the rate of urea hydrolysis and ammonium conversion to ammonia gas, and consequently increase volatilization.

The volatilization of surface urea increases linearly with increasing soil water content, until the soil reaches saturation. 0.1 inch of rain or irrigation dissolves fertilizer allows volatilization to occur while 0.5 inch of rain or irrigation pushes dissolved fertilizers about 2 inches deep into the soil and stopped the volatilization from occur if it is done within 2 days from fertilization.

Residues of crops left in the soil also increase the volatilization rate where the urease enzyme required for hydrolysis is produced by microorganisms that are 40 times more active in surface residues than in mineral soil. Besides, the residue of crops often exceeds the pH of the soil, which increases ammonia, impedes the movement of Nitrogen into the soil and also maintains soil moisture.

Volatile  $\text{NH}_3$  losses from irrigation and floodwater can be substantial. High  $\text{NH}_3$  concentration, high water pH, warm temperature and high wind speed all contribute to the probability of loss. Urea is more susceptible to volatilization when dispersed into flood water (such as flooded rice) than a fertilizer such as ammonium sulfate since the pH is likely to rise as urea is hydrolyzed to ammonium carbonate. Photosynthesis of plants and algae growing in water increases the pH of the water when flooded rice grows as  $\text{CO}_2$  is depleted during the daytime, causing pH to rise to 9.  $\text{CO}_2$  during plant respiration (forming carbonic acid) is released during the night and the water pH decreases again.

## **2.5 Ammonia volatilization problems to health and environment**

### **2.5.1 Issues on environment**

China has been the world's largest source of  $\text{NH}_3$  emissions, which is why they have experienced extreme severe and persistent haziness in recent years and has caused death up to five life years averagely. (Wu, Y., Gu, B., Erisman, J. W., Reis, S.,

Fang, Y., Lu, X., & Zhang, X., 2016)  $\text{NH}_3$  contributes significantly to the formation of secondary inorganic aerosols (SIA) resulting in alkaline gas pollution from the atmosphere. The finding shows that  $\text{NH}_3$  emissions have a significant impact on the average monthly (mean daily concentration)  $\text{PM}_{2.5}$  (fine particles with aerodynamic diameters below or equal to  $2.5 \mu\text{m}$ ), concentrations across major cities in China. (Zhao, M., Wang, S., Tan, J., Hua, Y., Wu, D., & Hao, J., 2016) They also suggested that  $\text{NH}_3$  is the critical pollutant of  $\text{PM}_{2.5}$  nucleation - related acid - base reactions.

Atmospheric  $\text{NH}_3$  also affects  $\text{NH}_4^+$  when the  $\text{NH}_3$  deposits on plant surfaces, dissolves and is washed into the soil where it can increase soil acidity and interfere with basic cation absorption (Pearson and Stewart 1993, Fangmeier et al. 1994, Krupa 2003). Negative effects on vegetation occur through direct toxicity when absorption exceeds the capacity of detoxification and, via N build-up. Excess ammonia will contribute to increased nitrification and denitrification, greenhouse gas emissions and eutrophication (deprivation of oxygen).

### 2.5.2 Issues on health

The ammonia gas that has volatilized also affects humans. Accidental acute exposure to ammonia makes eye injuries the most common cause of permanent disability. (Helmert et al., 1971; Jarudi and Golden, 1973). At a concentration of 140 ppm, gas ammonia is slightly irritating to the human eye and immediately irritates at 700 ppm (National Research Council, 1977; Griffiths and Megson, 1984). While exposure to ammonia vapor of 40 ppm led to headache, nausea and decreased appetite (National Research Council, 1977). Mild ammonia gas exposure usually leads to temporary, irritating symptoms of the respiratory system. Dyspnea, wheezing,

coughing, nasal discharge and bronchial secretions are immediate signs of accidental ammonia inhalation.

## **2.6 Methods for determining the volatilization of ammonia**

There are many methods of determining ammonia volatilization in farmland includes direct method and indirect methods.

### **2.6.1 Direct methods**

The direct methods are enclosure method where the soil, fertilizer and plants are enclosed together in an airtight container, and absorb the ammonia from the system by acid or alkaline substance, and then make quantitative determination. It is directly through the ammonia capture device to capture ammonia from the soil surface, commonly used in cell experiments, with advantages of simple equipment, easy to operate, good mobility, multi-point simultaneous measurement and spatial distribution measurement and high sensitivity (Song Y. S., Fan X. H., 2003) . But the ammonia in the sealed state of the process is completely different from the natural state.

Another direct method found was venting method where it was an improved version from enclosure method where it overcomes problem in the previous method. Compared with other methods, this method is simple, convenient combination, low sampling device was also made of polyvinyl chloride rigid plastic bucket (15 cm in diameter, 10 cm in height). Two sponges with a thickness of 2 cm and a diameter of 16 cm were uniformly impregnated with 15 mL of glycerol solution (50 mL of phosphoric acid + 40 mL of glycerol, constant volume to 1000 mL) and placed in a hard plastic bucket. The lower sponge is 5 cm from the bottom of the bucket and the upper sponge is flat with the top of the bucket. Upper sponges absorb ammonia in the air and prevent it from entering the lower sponge is absorbed within the device (Yang

W., Zhu A., Zhang J., et al., 2013 )

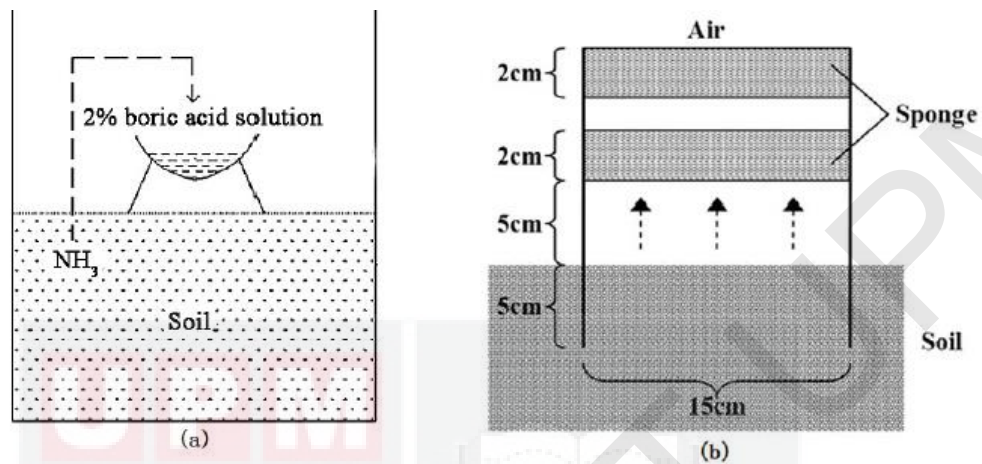


Figure 3 : Enclosure(a) and Venting(b) methods. (Yang W, Zhu A, Zhang J, et al., 2013)

### 2.6.2 Wind tunnel method

For indirect method, there are wind tunnel method where the volatilized ammonia gas sample were measured by calculating the difference between inlet and outlet of Teflon pipe. The remaining gas is expelled by the air pump and the solution in the bottle is determined by colorimetric method (Miola E. C., Rochette P., Chantigny M. H., et al., 2014). Finally, the ammonia volatilization per unit soil area at a given period of time was calculated by the difference between the outlet and the air inlet. However, this method requires expensive equipment.

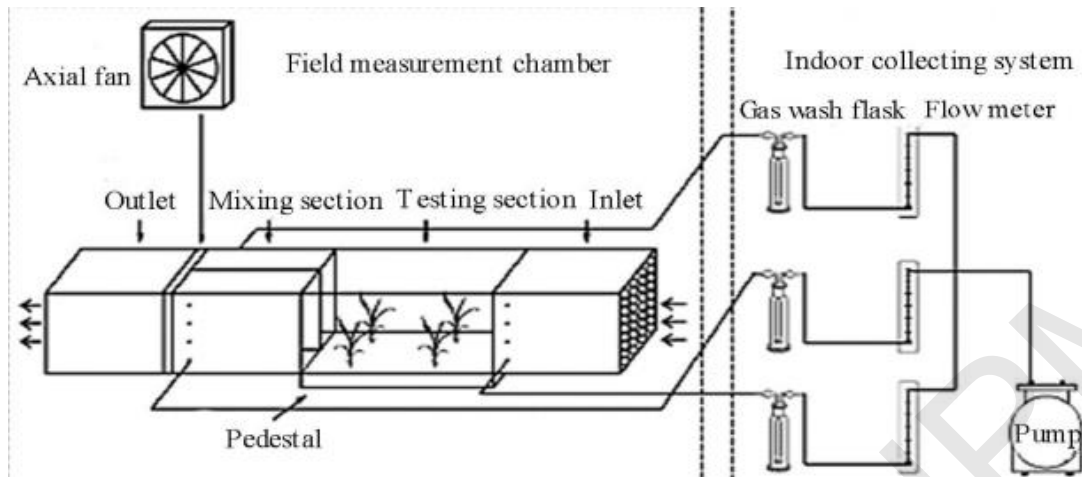


Figure 4 : Wind tunnel method for determination of ammonia volatilization. (Yang, J. et al., 2018)

### 2.6.3 Horizontal flux method

An integrated horizontal flux method was used to measure  $\text{NH}_3$  volatilization. The samplers were fixed on the vertical direction, and could freely rotate with the main wind direction to ensure that the most amount of ammonia in this direction could enter the samplers.

Previous researchers (Thompson et al., 1990; Sommer et al., 1991) stated that there was a significant relationship between the  $\text{NH}_3$  flux and the wind speed when the wind speed was lower than 2.5 m/s after slurry manure was applied. The result shows that in the daytime, the flux of ammonia volatilization was higher than that in the nighttime. This result indicates that the volatilization of  $\text{NH}_3$  is positively related to the air temperature. Ammonia volatilization also occurs due to strong sunshine, high temperature, the use of high pH fertilizer, or sometimes, the unsuitable fertilizer and slurry application methods. In this experiment, the pH of the slurry was 7.84, and the slurry was only applied on the surface of the soil.

#### 2.6.4 High-resolution measurement of ammonia emissions

The online measurement system consisted of three parts, including a wind tunnel system, a rotary gas collector (RGC) system and an ion chromatography. The wind tunnel system was used to simulate a real cultivation environment, which automatically sampled the airflow and kept the inside and outside micro-meteorological factors (e.g. sunlight, temperature, wind, and soil) basically the same (Braschkat et al., 1993; Mannheim et al., 1994; Huang et al., 2006).

In the RGC system,  $\text{NH}_3$  was absorbed and converted to ammonium by a chemical reaction with methanesulfonic acid. The RGC system was exploited to continuously collect gas online with high collection efficiency (Dong and Zeng, 2007). RGC System components included a worktable, sampling pipes, a rotary wet diffusion tube, flow controllers, inlet and outlet devices, analytical devices, and a power pump. The inlet of the diffusion tube was connected to the sampling pipe, while its outlet was attached to the flow controller; the outlet of the flow controller was connected to the pump. The samples were then transferred to glass bottles and analyzed by an ICS-900 (Dionex Corp., Sunnyvale, CA, USA) which was equipped with a CSRS 300 column (4 mm) for analysis of cations. The time resolution of the on-line measurement system was set at 30 min. The online  $\text{NH}_3$  emission monitoring system was proved to be stable and accurate in the current studies.

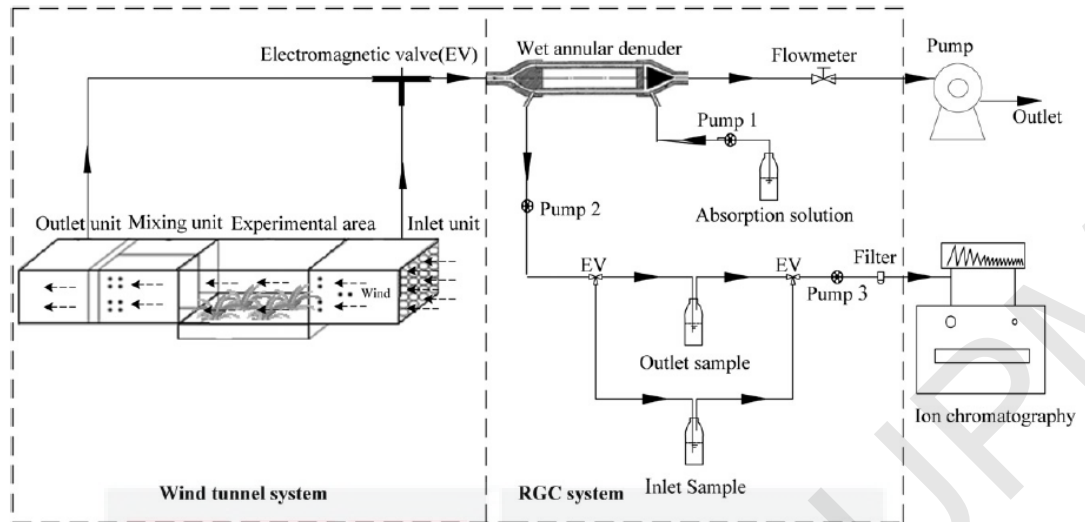


Figure 5 : Diagram of the ammonia emission on-line measurement system (Gong et al., 2013)

## CHAPTER 3

### MATERIALS AND METHOD

#### 3.1 Study Procedure

This study's main focus is on designing a device for measuring ammonia volatilization in Malaysia's flooded rice system. The first step is to study existing methods and devices for measuring ammonia volatilization and to compare the methods' advantages and disadvantages. The second is designing a cheap and suitable method for measuring ammonia volatilization and the third step is the development and evaluation by experiments of the designed methods.

At the beginning of this study, there are several methods of collecting ammonia gas in the air which are enclosure and micrometeorological method. The meteorological method was chosen for this study, where in free air above the surface of floodwater in rice fields, the vertical flux density of the gas is measured. Micrometeorological method was chosen because it tests the air above the surface of flood water and does not alter the natural environment, thus, it does not disrupt the mechanism of ammonia volatilization.

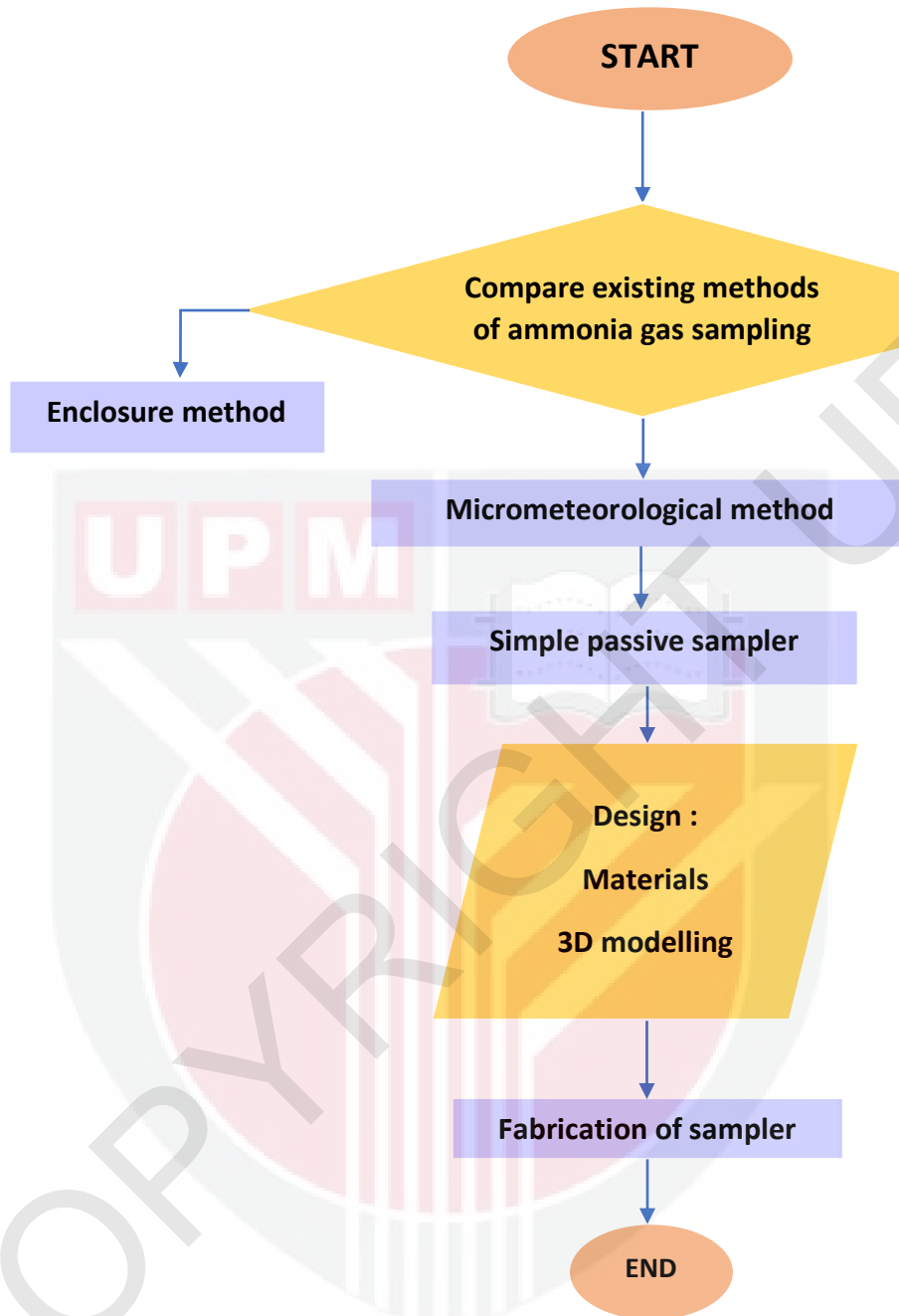


Figure 6 : Procedure of designing ammonia gas sampler device

## 3.2 Available micrometeorological sampler for gas collecting

### 3.2.1 Zurcher passive sampler

Zurcher sampler is a polypropylene vessel on the upper side with a lid and a Teflon gasket and an outlet on the lower side. The 0.2 mm thick Teflon membrane is positioned 7 mm underneath the tube edge. It isolates the surrounding air from strongly diluted HCl solution as absorption liquid. A custom protection box was built at the passive sampler's entrance to overcome air turbulence (Thöni, Seitler, Blatter, & Neftel, 2003).

The box comprises of an internal cylinder, opening at the base, enclosed by a closed container, at the top of the side wall with 21 air vents (diameter 9 mm) and seven at the bottom.

Ammonia disperses into an acid absorption solution via a permeable membrane. The primary advantages are: first, due to the absorption of ammonium-containing gases, it eliminates as much interference as possible, and second, it prevents the extraction phase of the covered surfaces in the samplers.

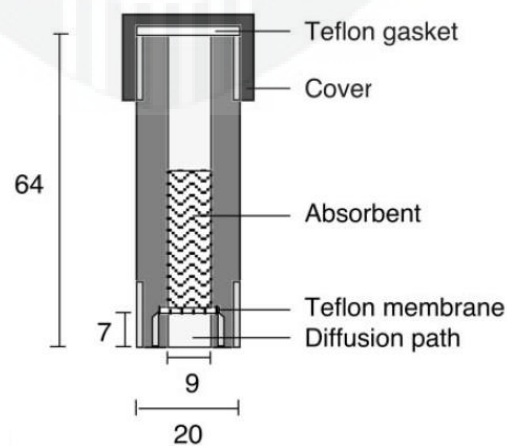


Figure 7 : Zurcher sampler with dimensions in mm

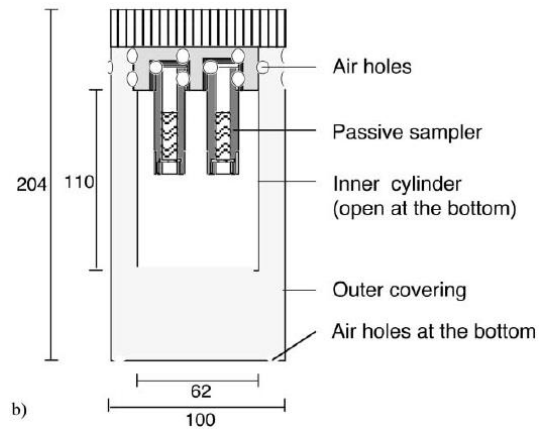


Figure 8 : Custom protection box for Zurcher sampler

### 3.2.2 Leuning ammonia sampler

This Leuning sampler comprises of an entry nozzle, an exterior cylinder equipped with pivots and fins to maintain the equipment parallel with the wind and a Venturi cover that supports a thin stainless steel orifice plate. Stainless steel was used to build inner absorbing structures consisting of a spiral of thin layer coiled on two wire retainers connected to a core steel tube.

Oxalic acid has been selected as the absorbent for ammonia since Ferm (1979) proved that ammonia gas could be fully absorbed by oxalic acid coatings at all relative air flow, humidity and that differences in temperature are not anticipated to impact the effectiveness of absorption.

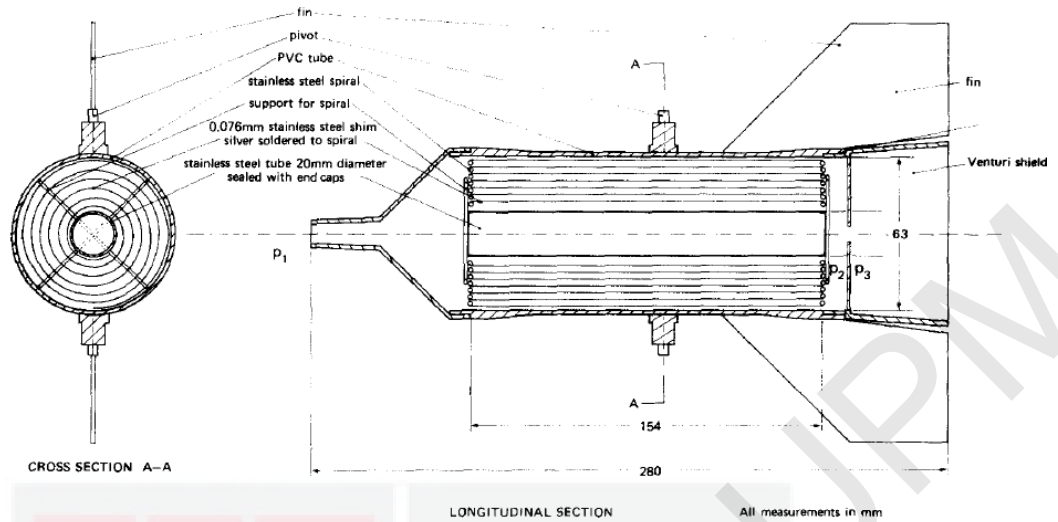


Figure 9 : Leuning sampler schematic diagram (R. Leuning, J. R. Freney, 1985)

### 3.2.3 Passive flux sampler

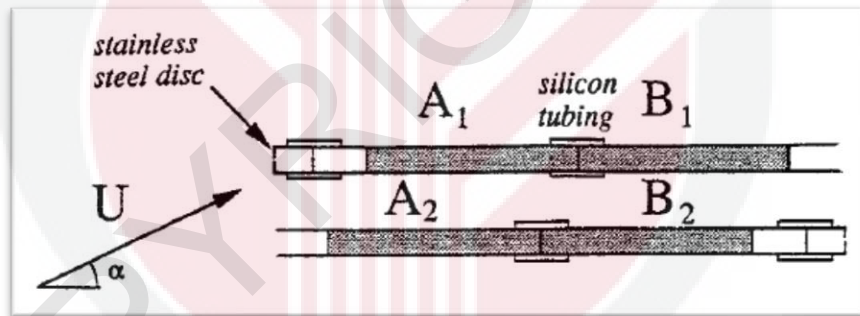


Figure 10 : Simple passive sampler

The sampler is made up of 2 glass vials, each 10 cm long and 0.7 cm in diameter. The two glass vials are connected to each other with silicon tubing and another pair are placed in series. Then, the sampler are placed at support pole at different heights to obtain different value of NH<sub>3</sub> emissions. A stainless steel disc with a hole was glued to the opening of sampler. The function of the stainless steel disc was to reduce the air velocity inside the pipes to obtain low friction force and maximum

sampling effectiveness of  $\text{NH}_3$ . Oxalic acid are used to coat inside the glass vials (J. K. Schjoerring, S. G. Sommer, 1991)

### 3.2.4 Ogawa sampler

The passive Ogawa sampler made by Ogawa USA, Pompano Beach, Florida is a handy tool for inspecting concentrations of atmospheric ammonia gas. The sampler consists of an air inlet on both side of a plastic casing. More than one gas can be sampled at once, as the two inlets are separated by the body's solid segment.

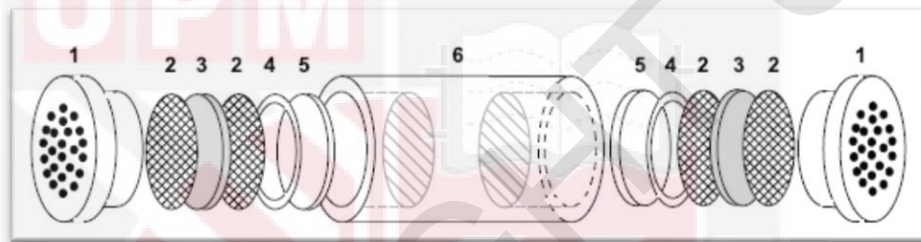


Figure 11 : Ogawa sampler (Roadman et al., 2003)

In the Ogawa sampler, from (1) is the end cap, (2) stainless steel screen sandwiching (3) filter made of glass fiber to gauze impurities, (4) retainer ring and (5) base pad. In its double-inlet setup made of Teflon with 2 cm outer diameter and 3 cm length, the sampler comprises two gas collection pads. The pads have a diameter of 14.5 mm and are specifically coated with a selected gas according to the type of gas needed to be sampled. To determine ammonia gas, the pads are soaked into citric acid solutions.

## CHAPTER 4

### RESULTS AND DISCUSSION

#### 4.1 Design of sampler

During this study, the design of readily available methods for determining ammonia volatilization are being analysed. Sampler from R. Leuning will be used as an example and comparison. The sampler performance has theoretically predicted and tested in wind tunnel, laboratory and also in field. The results shown that the air flow through the sampler is directly proportional to air flow externally and the ammonia that has absorbed by the sampler can be determined quantitatively.

##### 4.1.1 Conceptual design

The sampler works with inner oxalic acid coating to catch any ammonia in the atmosphere that flows through the sampler. This sampler has been fitted with nozzles and fins to maintain it always upwind.

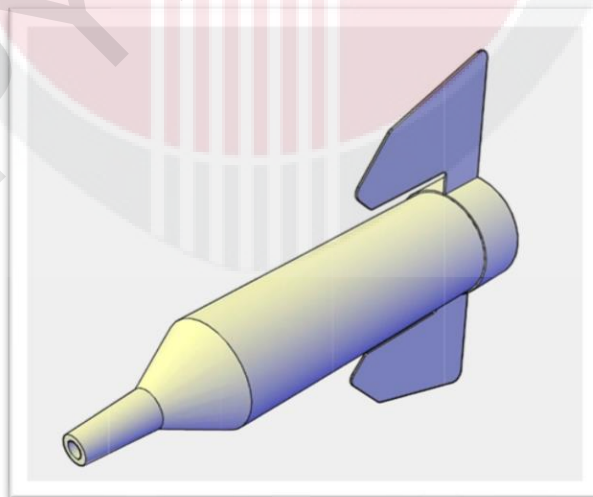


Figure 12 : 3D drawing of sampler

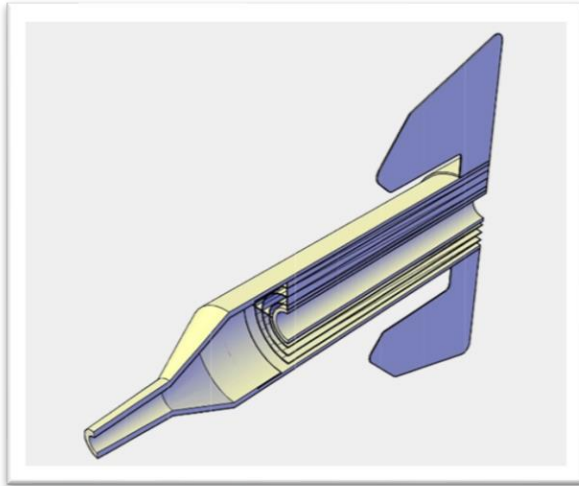


Figure 13 : Sectional drawing of the sampler



Figure 14 : Fabricated sampler

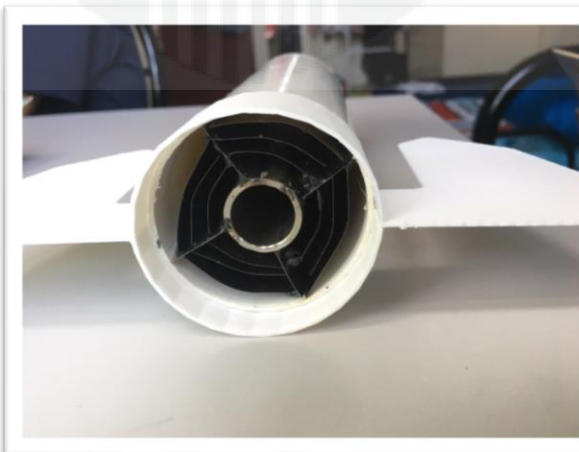


Figure 15 : Stainless steel assemble inside sampler body

The air enter nozzle at high speed and flows into the sampler. At the beginning, the air flows in turbulence. The nozzle decreases air speed inside the tubes to achieve a high  $\text{NH}_3$  collection efficiency. At the stainless steel assembly, air flow perpendicularly until it exits at the end of the sampler body. The fin is fitted to the sampler body to guide it along with the wind directions.

#### 4.1.2 Dimensions of sampler

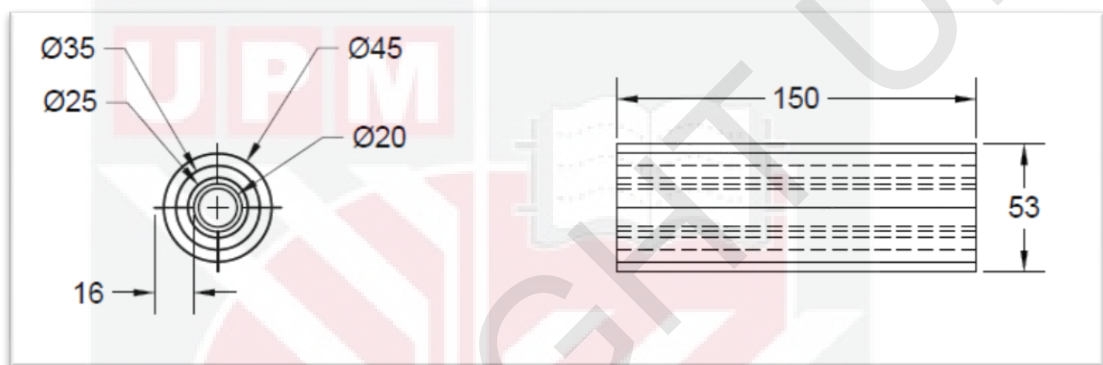


Figure 16 : Dimensions of stainless steel assembly (in mm)

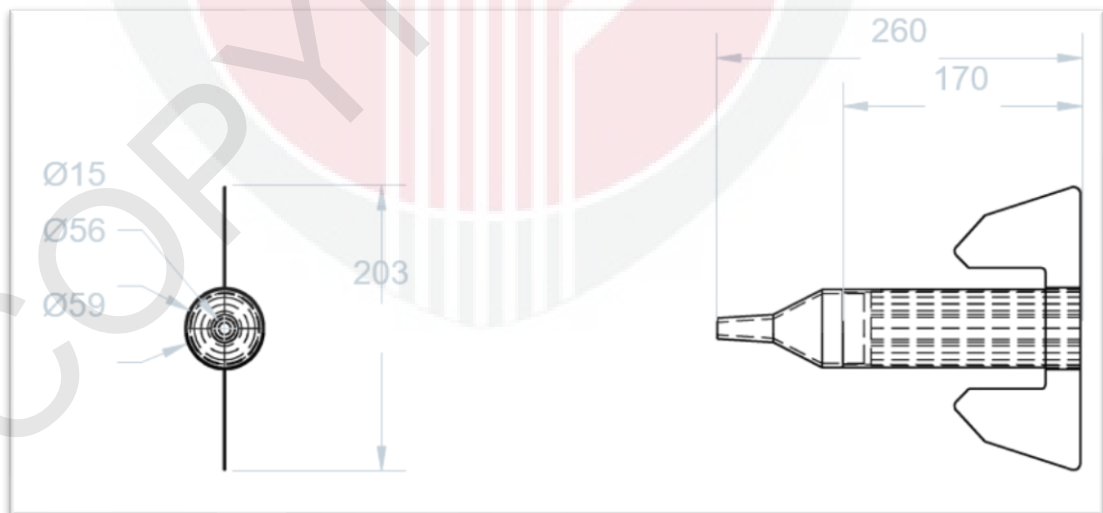


Figure 17 : Dimensions of sampler (in mm)

The outer part of the sampler that is made of PVC are 170 mm long, 56 mm diameter. After funnel is attached, the total length of the sampler is 260 mm. The opening of the funnel is 15 mm in diameter.

Meanwhile for the stainless steel assembly, the core rod have 20 mm diameter, coiled with 25 mm, 35 mm and 45 mm of stainless steel shim glued along four stainless steel shim attached on the surface of the rod.

#### 4.1.3 Materials

Stainless steel was used to construct the inner surfaces of sampler. The inner surface of the sampler acts as absorbent surface for  $\text{NH}_3$  gas. Oxalic acid is going to be used to coat the surfaces and stainless steel can withstand the chemical.



Figure 18 : Stainless steel shim 0.05 mm and 0.1 mm

PVC material is chosen as the sampler body because it is a great low-cost mechanical and chemical resistance characteristics. The workability of PVC with oxalic acid is good at both low (20°C) and high (60°C) temperature (“Chemical Resistance Guide,” 2003).

Epoxy has a good resistance towards oxalic acid. The electrical, thermal and chemical strength of epoxy resin is exceptional. Epoxy resilience with fibrous reinforcement or mineral fillers is commonly increased. In this design, Hardex 3 ton steelweld epoxy compound hardener & resin is used to fasten the stainless steel shims together with the rod. The parts are joined together and held for 30 minutes before it dries up strongly. Other methods of joining have been tested such as MIG welding and soldering, are not suitable because the method needs two parts that is going to be joined together, melted. This is because the stainless steel shim is really thin and not possible to melt before it is joined together.



Figure 19 : Epoxy resin and hardener

3D printing is used to create the fin and nozzle of the sampler. The material for both parts are Acrylonitrile Butadiene Styrene (ABS). ABS material has excellent physical characteristics, such as good tensile resistance and rigidity where the material are usually used to make car bumpers and LEGO products. It is also have a high resistance towards heat and chemical. ABS also have proven has an excellent resistance to oxalic acid (Röchling Formaterm AB)

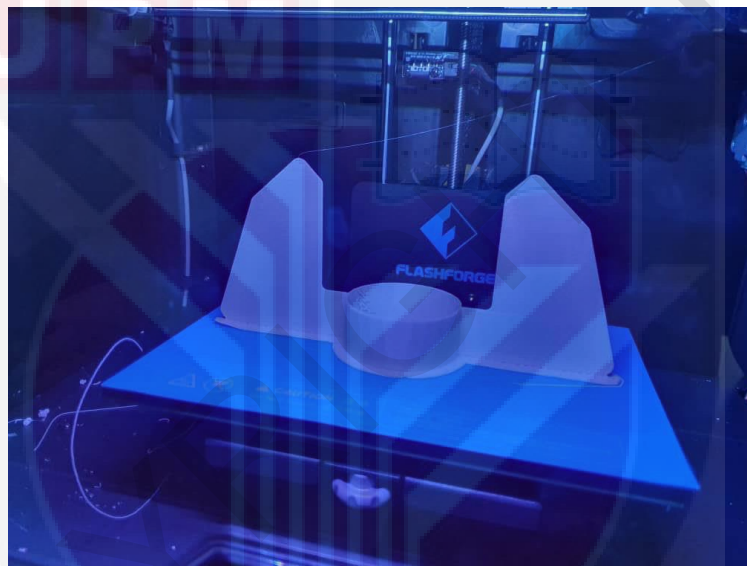


Figure 20 : Process of making fin and the fittings

## CHAPTER 5

### CONCLUSION AND RECOMMENDATIONS

#### 5.1 Conclusion

The sampler has the practical advantages of being easy to deploy, no climate parameters need to be quantified, no electricity is required, and the experimental area's surroundings are not specifically restricted.

With the design of ammonia gas sampler, ammonia losses in flooded rice systems can be analysed. It can be further controlled by suitable methods such as stabilizing the pH of floodwater, placing fertilizers at the right time of the day, using slow release fertilizers and deep placement of fertilizers instead of surface application.

Last but not least, the loss of ammonia can be effectively and inexpensively assessed without the need for a big workforce.

#### 5.2 Recommendations

The recommendations that can be suggested for this study in the future is the researcher needs to set up a laboratory testing to determine this devices' efficiency in trapping ammonia gas. The inner part of the sampler also needs to be reviewed to increase the surface area and choosing materials that can keep it light-weighted so that the sampler can always be in the direction of wind.

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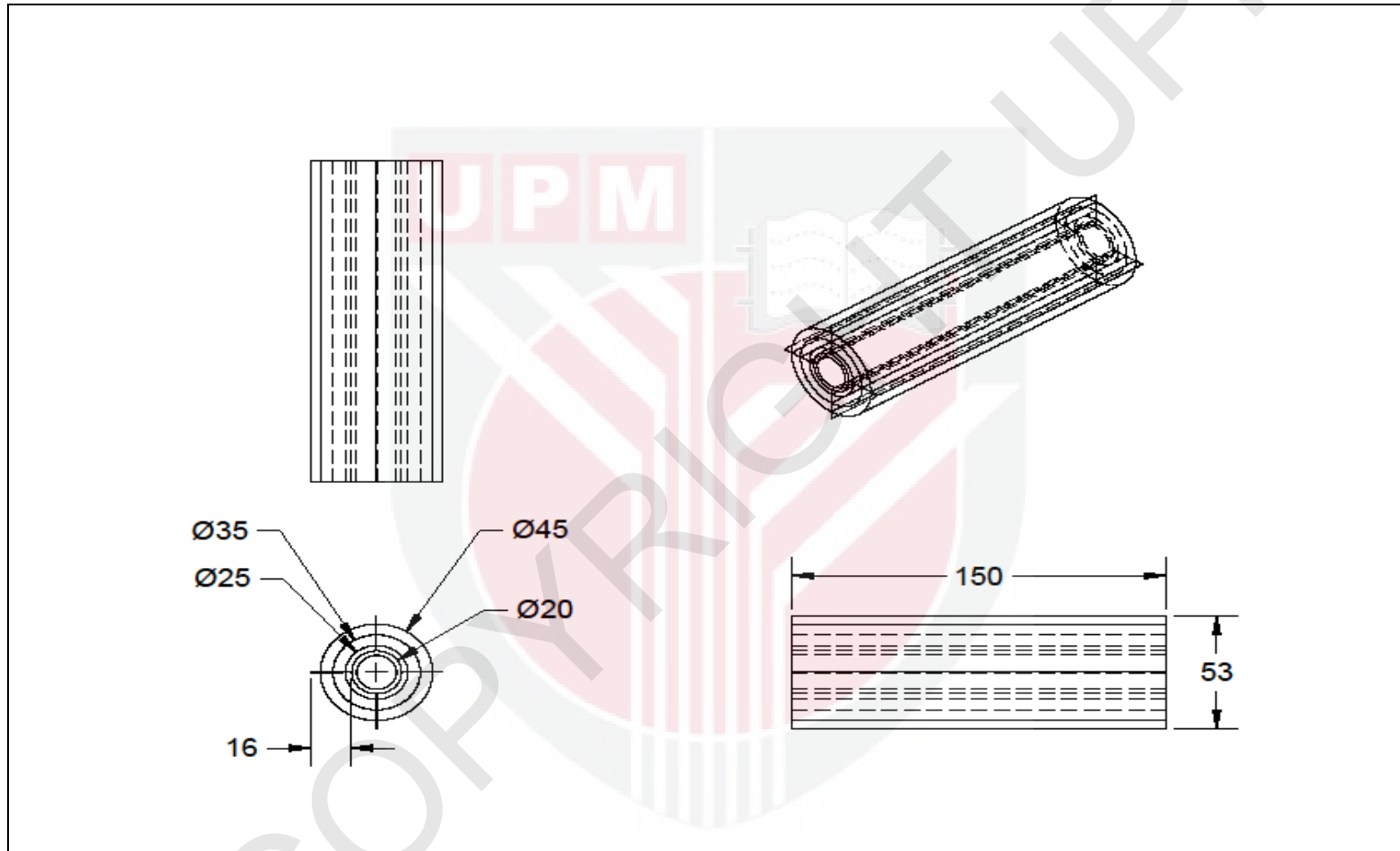
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## APPENDICES



Appendix 1 : Orthographic drawing of stainless steel assembly (dimensions are in mm)

